**The Mole Guided Notes**

* A mole is a form of \_\_\_\_\_\_\_\_\_\_\_\_\_\_ and is used to count the number of \_\_\_\_\_\_\_\_\_ in a \_\_\_\_\_\_\_\_\_\_\_\_
* 1 mole is equal to \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ just like \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ is equal to 12 cookies
* How many cookies/atoms are in 5 dozen?
* How many atoms are in 5 moles?
* Write the 2 conversion factors for the mole below:
* Moles ↔ Atoms
  + If atoms is given and you want moles…
  + If moles is given and you want atoms…
* How many moles of Mg are in 1.25 x 1023 atoms?
  + Need to go from \_\_\_\_\_\_\_\_\_\_\_\_\_\_ to \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
  + Write the example below:
* How many atoms of C are in 2.0 moles?
  + Need to go from \_\_\_\_\_\_\_\_\_\_\_\_\_\_ to \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
  + Write the example below:
* Molar mass is the \_\_\_\_\_\_\_\_\_\_ in grams of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of the substance
* Calculate the molar mass of HCl below:
* Moles ↔ Mass (g)
  + If grams is given and you want moles…
  + If moles is given and you want grams…
* How many moles are in 92.2 grams of Fe?
  + Need to go from \_\_\_\_\_\_\_\_\_ to \_\_\_\_\_\_\_\_\_\_
  + The molar mass of Fe is \_\_\_\_\_\_\_\_\_\_\_
  + Write the example below:
* What is the mass of 9.45 moles of HCl?
  + Need to go from \_\_\_\_\_\_\_\_\_ to \_\_\_\_\_\_\_\_\_\_
  + The molar mass of HCl is \_\_\_\_\_\_\_\_\_\_\_
  + Write the example below: